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comprehensive chemistry according to the new syllabus prescribed by central board of secondary education cbse

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a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level lowest energy of the particle chapter 2 thermodynamics i brief resume of first and second law of thermodynamics entropy changes in reversible and irreversible processes variation of entropy with temperature pressure and volume entropy concept as a measure of unavailable energy and criteria for the spontaneity of reaction free energy enthalpy functions and their significance criteria for spontaneity of a process partial molar quantities free energy volume heat concept gibb s duhem equation chapter 3 chemical dynamics i effect of temperature on reaction rates rate law for opposing reactions of ist order and iind order rate law for consecutive parallel reactions of ist order reactions collision theory of reaction rates and its limitations steric factor activated complex theory ionic reactions single and double sphere models influence of solvent and ionic strength the comparison of collision and activated complex theory chapter 4 electrochemistry i ion ion interactions the debye huckel theory of ion ion interactions potential and excess charge density as a function of distance from the central ion debye huckel reciprocal length ionic cloud and its contribution to the total potential debye huckel limiting law of activity coefficients and its limitations ion size effect on potential ion size parameter and the theoretical mean activity coefficient in the case of ionic clouds with finite sized ions debye huckel onsager treatment for aqueous solutions and its limitations debye huckel onsager theory for non aqueous solutions the solvent effect on the mobality at infinite dilution equivalent conductivity 2 vs concentration c 1 2 as a function of the solvent effect of ion association upon conductivity debye huckel bjerrum equation chapter 5 quantum mechanics ii schrodinger wave equation for a particle in a three dimensional box the concept of degeneracy among energy levels for a particle in three dimensional box schrodinger wave equation for a linear harmonic oscillator its solution by polynomial method zero point energy of a particle possessing harmonic motion and its consequence schrodinger wave equation for three dimensional rigid rotator energy of rigid rotator space quantization schrodinger wave equation for hydrogen atom separation of variable in polar spherical coordinates and its solution principle azimuthal and magnetic quantum numbers and the magnitude of their values probability distribution function radial distribution function shape of atomic orbitals s p d chapter 6 thermodynamics ii classius clayperon equation law of mass action and its thermodynamic derivation third law of thermodynamics nernest heat theorem determination of absolute entropy unattainability of absolute zero and its limitation phase diagram for two completely miscible components systems eutectic systems calculation of eutectic point systems forming 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