

Experiment 34 An Equilibrium Constant Answers

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the present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields the pedagogical strategy i have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature the procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions there is repeated emphasis on the use of corrected

conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations the coverage of solution equilibrium phenomena is not however exhaustively comprehensive rather i have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

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a straightforward explicit derivation of the equilibrium constant is presented together with a discussion of the dimensional relationships that exist among equilibrium constants commonly used author

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