

Chemfax Acid Base Titrations Lab Prelab

Answers

Chemfax Acid Base Titrations Lab Prelab Answers Chemfax AcidBase Titrations Lab PreLab Answers This document provides answers to the prelab questions for a Chemfax experiment on acid base titrations It aims to guide you through the fundamental concepts and calculations involved in the experiment

1 Define the following terms

a Titration Titration is a quantitative analytical technique used to determine the concentration of an unknown solution called the analyte by reacting it with a solution of known concentration called the titrant The titrant is carefully added to the analyte until the reaction is complete which is indicated by a change in the solutions color or pH

b Titrant The titrant is a solution of known concentration used in a titration It reacts with the analyte which is the solution of unknown concentration

c Analyte The analyte is the solution of unknown concentration that is being analyzed in a titration It reacts with the titrant to determine its concentration

d Equivalence Point The equivalence point in a titration is the point at which the moles of titrant added are stoichiometrically equal to the moles of analyte present in the solution At the equivalence point the reaction between the titrant and analyte is complete

e End Point The endpoint in a titration is the point at which a visual indicator signals that the reaction is complete The endpoint is usually very close to the equivalence point but not always identical

2 Explain the difference between a strong acid and a weak acid

Strong Acids Completely ionize in solution meaning they donate all of their protons H^+ to the solvent Have very low pH values less than 1

2 Examples Hydrochloric acid HCl Sulfuric acid H_2SO_4 Nitric acid HNO_3

Weak Acids Partially ionize in solution meaning they only donate a fraction of their protons Have higher pH values than strong acids usually between 2 and 6

Examples Acetic acid CH_3COOH Carbonic acid H_2CO_3 Phosphoric acid H_3PO_4

3 Explain the difference between a strong base and a weak base

Strong Bases Completely dissociate in solution releasing all of their hydroxide ions OH^- into the solution Have very high pH values greater than 13

Examples Sodium hydroxide $NaOH$ Potassium hydroxide KOH Calcium hydroxide $Ca(OH)_2$

Weak Bases Partially dissociate in solution meaning they only release a fraction of their hydroxide ions Have lower pH values than strong bases usually between 8 and 12

Examples Ammonia NH_3 Methylamine CH_3NH_2 Pyridine C_5H_5N

4 What is a pH indicator How does it work

A pH indicator is a substance that changes color in response to changes in pH It is used to visually signal the endpoint of a titration which is often very close to the equivalence point

How it works pH indicators are typically weak acids or bases that exhibit different colors in their acidic and basic forms When an indicator is added to a solution it exists in equilibrium between its acidic HIn and basic In^- forms

H In Acidic Form Basic Form The color of the solution is determined by the relative concentrations of these two forms In acidic solutions the acidic form HIn predominates resulting in one color In basic solutions the basic form In predominates resulting in a different color As the pH of the solution changes the equilibrium shifts leading to a change in the color of the indicator

5 Describe the procedure for standardizing a solution of NaOH

Standardization of NaOH solution

3 1 Preparation Weigh out a known mass of a primary standard usually potassium hydrogen phthalate KHP KHP is a weak acid that is stable nonhygroscopic does not absorb water and has a high molar mass making it an ideal primary standard

2 Dissolution Dissolve the weighed KHP in distilled water and transfer it to a clean dry flask

3 Titration Add a few drops of phenolphthalein indicator to the KHP solution This indicator will turn pink in the presence of excess NaOH Slowly titrate the KHP solution with the NaOH solution until a faint pink color persists for at least 30 seconds

4 Calculation Record the volume of NaOH solution used to reach the endpoint Use the mass of KHP its molar mass and the volume of NaOH used to calculate the concentration of the NaOH solution

5 Repeat Repeat the titration at least twice to ensure accurate results

6 A solution of NaOH is standardized using 0.650 g of KHP The titration requires 28.40 mL of the NaOH solution to reach the endpoint Calculate the molarity of the NaOH solution

Calculations

1 Moles of KHP Molar mass of KHP KHCHO_2 204.22 g/mol Moles of KHP $0.650 \text{ g} / 204.22 \text{ g/mol} = 0.00318 \text{ mol}$

2 Molarity of NaOH Volume of NaOH solution 28.40 mL 0.02840 L Since the reaction between KHP and NaOH is a 1:1 mole ratio the moles of NaOH are equal to the moles of KHP 0.00318 mol Molarity of NaOH $0.00318 \text{ mol} / 0.02840 \text{ L} = 0.112 \text{ M}$ Therefore the molarity of the NaOH solution is 0.112 M

7 A 25.00 mL sample of an unknown acid is titrated with 0.100 M NaOH The titration requires 35.40 mL of NaOH solution to reach the endpoint Calculate the molarity of the acid

Calculations

1 Moles of NaOH Molarity of NaOH 0.100 M Volume of NaOH 35.40 mL 0.03540 L Moles of NaOH $0.100 \text{ mol/L} \times 0.03540 \text{ L} = 0.00354 \text{ mol}$

2 Moles of Acid Assuming the acid is monoprotic donates one proton the mole ratio between NaOH and the acid is 1:1 Therefore the moles of acid are also 0.00354 mol

3 Molarity of Acid Volume of acid 25.00 mL 0.02500 L Molarity of acid $0.00354 \text{ mol} / 0.02500 \text{ L} = 0.142 \text{ M}$ Therefore the molarity of the unknown acid is 0.142 M

8 What are some of the sources of error in an acidbase titration

Sources of error in acidbase titrations

- Inaccurate measurement of titrant volume This can occur due to misreading the burette or air bubbles in the burette tip
- Inaccurate measurement of analyte volume This can occur due to misreading the pipette or volumetric flask
- Inaccurate endpoint determination This can occur if the indicator used is not a good match for the titration or if the endpoint is not observed carefully
- Improper standardization of the titrant If the titrant is not accurately standardized the results of the titration will be inaccurate
- Contamination of reagents If the titrant or analyte solutions are contaminated the results of the titration will be inaccurate
- Temperature variations The concentration of solutions can vary with temperature which can affect the accuracy of the titration
- Incomplete reaction If the

reaction between the titrant and analyte is not complete the results of the titration will be inaccurate Conclusion This document provided answers to the prelab questions for a Chemfax experiment on acid base titrations Understanding the fundamental definitions concepts and calculations involved is crucial for conducting successful and accurate titrations

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coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules this book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry it not only

highlights the key fundamental components of the coordination chemistry field it also exemplifies the historical development of concepts in the field in order to graduate as a chemistry major that fills the requirements of the american chemical society a student needs to take a laboratory course in inorganic chemistry most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine the environment molecular biology organic synthesis and inorganic materials

this laboratory manual is intended for a two semester general chemistry course the procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life this lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes by the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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this manual includes all classic biochemistry techniques such as hplc or enzyme kinetics and is complete with numerous problem sets relating to each topic

this e book is a collection of exercises designed for students studying chemistry courses at a high school or undergraduate level the e book contains 24 chapters each containing various activities employing applications such as ms excel spreadsheets and spartan computational modeling each project is explained in a simple easy to understand manner the content within this book is suitable as a guide for both teachers and students and each chapter is supplemented with practice guidelines and exercises computer based projects for a chemistry curriculum therefore serves to bring computer based learning a much needed addition in line with modern educational trends to the chemistry classroom

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